

## Fabrication and Characterization of Nano Titania-Magnetic Reduced Graphene Oxide with Highly Photocatalytic Degradation of Tartrazine Dye

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**Abstract:** In this paper, magnetic nanocomposites are synthesized by loading reduced graphene oxide (RG) with two components of nanoparticles consisting of titanium dioxide (TiO<sub>2</sub>) and magnetite (Fe<sub>3</sub>O<sub>4</sub>) with varying amounts. The resulting TiO<sub>2</sub>/magnetite reduced graphene oxide (MRGT) composite demonstrate intrinsic visible light photocatalytic activity, on degradation of tartrazine (TZ) dye from a synthetic aqueous solution. Specifically, it exhibits higher photoactivity than magnetite reduced graphene oxide (MRG) and TiO<sub>2</sub> nanoparticles. The photocatalytic degradation of (TZ) by MRGT composites under visible light irradiation was examined by varying the operational parameters such as photocatalyst amounts, organic dye concentrations, and irradiation time. The structural and magnetic features of the prepared composite photocatalysts were investigated by powder X-ray diffraction (XRD), Fourier transform infrared spectra (FT-IR), transmission electron microscopy (TEM), UV-vis diffuse reflectance spectra (UV-vis/DRS) -, Raman and vibrating sample magnetometer (VSM). The photocatalytic degradation of TZ dye when using MRG and TiO<sub>2</sub> for 3 h under visible light was 35% and 10 % respectively, Whereas for MRGT it was more than 95% . The higher photocatalytic efficiency of MRGT is due to the existence of reduced graphene oxide and magnetite which enhance the photocatalytic efficiency of the composite in visible light.

**Keywords:** Reduced Graphene oxide, Tartrazine, Titanium dioxide, Photocatalyst, GO, Magnetite.

### 1. Introduction

Water contamination has become one of the dominant problems in everyday life. Increasing the manufacturing wastewater from materials, foods, medicines, etc. which containing various kinds of dyes such as make it more worrying[1].

Strongly colored effluents can make environmental problems, in addition to influence, the human life as the greater part of them are cancer-causing and mutagenic[2]. The reports demonstrate that it is approximately 10–15% of the dyes is missed in the effluent during dyeing procedure[3].

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These dyes are considered as nonoxidizable by traditional physical and biological remediation because of the complex of their molecular structure and their large size molecules. So, their decolorization is considered as an essential procedure in wastewater treatment[4].

One of these dyes is tartrazine dye, which is a water soluble acidic azo dye (FD & C Yellow No. 5, EEC Number = E102, CI Number = 19140) which has the following chemical structure [trisodium 5-hydroxy-1-(4-sulfonatophenyl)-4-(4-sulfonatophenylazo)-H-pyrazol-3-carboxylate] [5, 6].

Tartrazine dye is a water-soluble powder which is greatly used in nutrition products, medicines, and cosmetics[7-9]. The suitable dose per person is 0–7.5mg/kg body weight[10]. Tartrazine is considered as a dangerous azo dye because it catalyzes ADHD syndrome, asthma, migraine and thyroid cancer, and other reactions can include blurred vision, itching, purpura, eczema, mutagenetics, and carcinogenetic [11].

At the present time, due to these problems, it is necessary to develop highly effective and sensitive techniques to remove pollutants from water environment. Various types of techniques were used for this purpose, for example, photocatalytic degradation, heterogeneous catalytic reaction, electrochemical removal, and different kinds of adsorbent such as adsorbents prepared from industrial waste, carbon nanotube composite and natural adsorbents, and bottom ash and de-oiled soya were utilized.

In recent times, heterogeneous photocatalytic oxidation, one of the Advanced Oxidation Processes (AOPs), has involved much attention and has confirmed to be an effective technology for degradation of an organic contaminant at ambient conditions[12]. Its advantage is the complete degradation of organic pollutants compared with the other water purification techniques[13].

In this technique, photogenerated charge carriers in a semiconductor photocatalyst and the creation of extremely reactive chemical species like hydroxyl radicals could mineralize a wide range of organic contaminants rapidly and non-selectively [14].

The most usually utilized photocatalyst is  $\text{TiO}_2$  which has been utilized as a promising photocatalyst for the degradation of organic molecules from wastewater due to its distinguished and distinctive properties like high stability, low toxicity and low cost[15]. However, the photocatalytic efficiency of  $\text{TiO}_2$  is restricted due to the high recombination rate of electron–hole within the  $\text{TiO}_2$  particles with respect to the time scale of the required redox reaction, and its wide band gap. So, only 4% to 5% of the sunlight (UV irradiation) can activate  $\text{TiO}_2$  at wavelengths smaller than 380 nm[16].

The method used to overcome the wide band gap of  $\text{TiO}_2$  is the modification of the electronic band structure of the  $\text{TiO}_2$  by chemical reaction with a substance which is used as a source of charge-carrier traps and increase the lifetime of separated electron–hole pairs. In this study, Reduced Graphene oxide (RG) and Magnetite (M) were used with Titanium dioxide (T) to form MRGT composite[17, 18].

Graphene (RG), a single-atom layer of graphite, has been displayed to own amazing catalytic, electrocatalytic, electrochemical sensing and electrochemical energy conversion properties[19]. This two-dimensional network of carbon with high theoretical specific surface area has an excellent mechanical strength, structural elasticity and chemical stability[20]. It also displays brilliant mobility of charge carriers and noble thermal conductivity. These

properties make graphene very desirable support for semiconductor and metal nanoparticles[21].

The incorporation of graphene (RG) into  $\text{TiO}_2$  may give good advantages as follows:

(a) Increased adsorptivity, where graphene displays a particularly worthy adsorptive property;  
(b) Extended visible light absorption, in which the chemical bonds of Ti-O-C provide a red shift in the photo-responding range that facilitate an extra efficient utilization of visible light for the catalyst. (c) Inhibit charge recombination [22].

Recently, magnetic separation technology has provided a suitable method for eliminating and recycling magnetic particles/composites by applying suitable magnetic field.  $\text{Fe}_3\text{O}_4$  (M) nanoparticles with a fast and efficient magnetic response can allow the separation of catalysts from reaction mixture with any problem and efficiently. Thus, MRGT composite is needed to be recoverable and an efficient photocatalyst towards contaminant degradation [23].

In this work, the MRGT composite photocatalysts were synthesized by various methods. The photocatalytic activity of the MRGT composite photocatalysts was exploited for the photodegradation of a model organic dye tartrazine (TZ) in a suspension system. The main target of our work is to estimate the potential application of this kind of material in the removal of organic contaminants in water [24].

## 2. Experimental

### 2.1 Materials

Graphite powder 99.5%, Ferric chloride 99%, Ammonia solution 30%, titanium (IV) isopropoxide, and isopropyl alcohol were purchased from Sigma Aldrich. Sodium nitrate 99%, Potassium permanganate 98%, Sulfuric acid 98%, Hydrogen peroxide 35%, Ferrous Sulfate 98% were obtained from alpha chemicals, Egypt. Commercial industrial tartrazine were acquired from local markets, Egypt. All purchased compounds are used as received, with no further purification.

### 2.2 Synthesis of graphene oxide (GO)

Graphene oxide was synthesized by Hummer method using manufactured graphite powder as a starting material. The GO preparation is described by the following procedures: graphite (5 g), sodium nitrate ( $\text{NaNO}_3$ , 2.50 g) and concentrated sulfuric acid ( $\text{H}_2\text{SO}_4$ , 115ml) were added into a 2 L beaker, which is kept at 5 °C in an ice bath under stirring for 10 min. At that point, potassium permanganate ( $\text{KMnO}_4$ , 15 g) was gradually added into the beaker to inhibit any severe reaction that might take place. The mixture was kept under stirring at 5 °C for 1 h, after that it was left for 24 hrs at room temperature preserving an intermediate stirring.

Distilled water (230 ml) was then added to the mixture, which enforces the temperature to expand to 98 °C by the effect of the hydration heat. The mixture was kept at this temperature for 30 min under stirring. As a final stage, distilled water (700 ml) and hydrogen peroxide (30%, 300 ml) were added. The obtained product had a brown /yellowish color and washed three times with distilled water to remove any remnant ions in the solution. Finally, GO was isolated by vacuum filtration from the solution and later it was dried in oven at 60 °C for 12 hrs [25].

### 2.3 Synthesis of magnetite reduced graphene oxide (MRG) composite

MRGO were prepared by in-situ chemical co-precipitation method. The procedure is as per the following: 1:2 molar ratio of ferrous sulfate ( $\text{FeSO}_4$ ) and ferric chloride ( $\text{FeCl}_3$ ) in 80ml of distilled water under  $\text{N}_2$  gas. The solution kept under mechanical stirring for 30 min. Then the solution was heated to  $70^\circ\text{C}$  and kept under stirring for another 30 min. Different amount ratio 10, 20 and 30 wt.% of GO relative to magnetite was suspended in 50 mL of distilled water by ultrasonic treatment for 15 min until it turned out to be clear without any observed particles to produced MRG10, MRG20 and MRG30 respectively. GO solutions were added to the mixture of iron solution and the whole mixture was under stirring for another 10 min then ammonia solution (30%) was added to the mixture until pH of the mixture is 10. Moreover of sodium borohydride was added to the solution under mechanical stirring for 10 min in ordered to make a reduction of GO nanoparticle to obtain reduced graphene oxide. The sample was separated by vacuum filtration from the solution and washed with distilled water and ethanol two times respectively, finally dried in the furnace at  $60^\circ\text{C}$  for 12h [26].

### 2.4 Synthesis of $\text{TiO}_2$ nanoparticles

$\text{TiO}_2$  nanoparticles were synthesized by hydrolysis method as the following procedure. 7ml of Titanium (IV) isopropoxide  $\text{Ti}[\text{OCH}(\text{CH}_3)_2]_4$  was dissolved in 100 ml of isopropyl alcohol ( $(\text{CH}_3)_2\text{CHOH}$ ) and the solution was under stirring for 1 h at room temperature. After that, 400 ml of distilled water was added abruptly to the mixture. The obtained solution was stirred for another 3 h at room temperature. The sample was filtrated by vacuum filtration from the solution and washed with distilled water and ethanol three times respectively. The obtained sample was dried at  $40^\circ\text{C}$  in the oven for 6 hr. The synthesis process was then finished by calcinating the precipitated particles at  $400^\circ\text{C}$  for 2 h in air [27].

### 2.5 Synthesis of $\text{TiO}_2$ /MRG composite

MRGT composites were prepared with a weight ratio between MRG 20 and  $\text{TiO}_2$  nanoparticles at (3:1, 1:1 and 1:3) to produced MRGT 25, MRGT 50 and MRGT 75 respectively. Where  $\text{TiO}_2$  nanoparticles and MRG 20 were suspended each one in 50ml of ethanol by ultrasonic treatment for 10 min until it turned out to be clear without any observed particles then the two solutions were mixed together by ultrasonic treatment for 10min. After that, the sample was separated by vacuum filtration from the solution and it was put in the oven at  $60^\circ\text{C}$ .

### 2.6 Characterization

The phases of the designed nanomaterials were examined via an X-ray diffractometer (XRD, Shimadzu XD-1) with a Cu Ka radiation at 40 kV and 30 mA over the  $2\theta$  range of  $4-80^\circ$  at a scanning speed of  $4^\circ/\text{min}$  with a sampling angle interval of  $0.04^\circ$ . Function groups and bonds are identified using FTIR spectrometer PerkinElmer (model spectrum one FTIR spectrometer, USA). Spectra were recorded between  $4000$  and  $500\text{ cm}^{-1}$  using standard KBr pellets. In addition to FTIR, Raman test was carried out using the dispersive Raman microscope (Model Sentera, Bruker, Germany) instrument at laser wave length 532 nm. Morphology and particle size were determined by high resolution transmission electron microscope (TEM-JEOL JEM 2100) and composition was analyzed by Energy dispersive X-ray spectroscope (EDX) associated with TEM device. Magnetic measurements were carried out at room temperature using a vibrating sample magnetometer (VSM, 735VSM, Model7410; Lake Shore, Westerville, Ohio, USA) with a maximum magnetic field of 31 kOe. UV-VIS/DR spectra were recorded using Jasco V 530 spectrometer (Japan) equipped with the integrating sphere accessory for diffuse reflectance spectra.

## 2.7 Photocatalytic activity

To check the photocatalytic activity of MRG and MRGT composites, a reaction of TZ solved in water is performed under simulated visible illumination of a halogen lamp (500W). Added catalyst to 50 mL aqueous TZ solution (50ppm) and continuously mechanical stirred for 30 min in dark box, until adsorption–desorption equilibrium is obtained between TZ dye and photocatalyst. The suspension is irradiated as a function of time. Samples are then withdrawn regularly from the reactor and the photocatalyst is separated by an external magnet. The clean transparent solution is analyzed by agilent technologies Cary 60 UV–vis spectroscopy.

## 3. Results and Discussion

### 3.1 Characterization of photocatalysts

Fig. (1.a) shows the diffraction peaks at  $2\theta$  around 30.27, 35.75, 43.6, 53.73, 57.5, and 63.1, can be listed to the (220), (311), (400), (422), (511) and (440) lattice planes of cubic  $\text{Fe}_3\text{O}_4$  (JCPDS card No. 88-0315), respectively. When the amount of  $\text{Fe}_3\text{O}_4$  diminished, its diffraction peaks in the drawing were slowly expanded and debilitated, because of little crystal size and more regrettable crystalline of the  $\text{Fe}_3\text{O}_4$  [28]. There might be two explanations behind no peaks from carbon were watched. Firstly, the existence of magnetite decreases the accumulation of graphene sheets, which makes more monolayer graphene prompting to less weak peak from carbon to be watched. Secondly, a strong signal of iron oxides has a tendency to dominate the weak carbon peaks [29].

Crystal structure of MRGTx samples where x is 25, 50 and 75 present in Fig. (1.b). The diffraction peaks at  $2\theta$  25.3°, 37.8°, 48°, 54°, 55.2° are attributed to (101), (004), (200), (105) and (211) crystallographic planes of anatase phases of  $\text{TiO}_2$  (JCPDS 21-1272). A (002) diffraction peak at 10.80° was not seen in  $\text{TiO}_2$ -RG composites, which demonstrates that the GO was decreased into RG. The reflection of RG is not seen in the XRD patterns of MRGT composites, perhaps because of the small amount of RG and low intensity of RG. In addition, the distinctive peak of RG at 24.5° might be covered by the fundamental peak of anatase  $\text{TiO}_2$  at 25.3 [30].

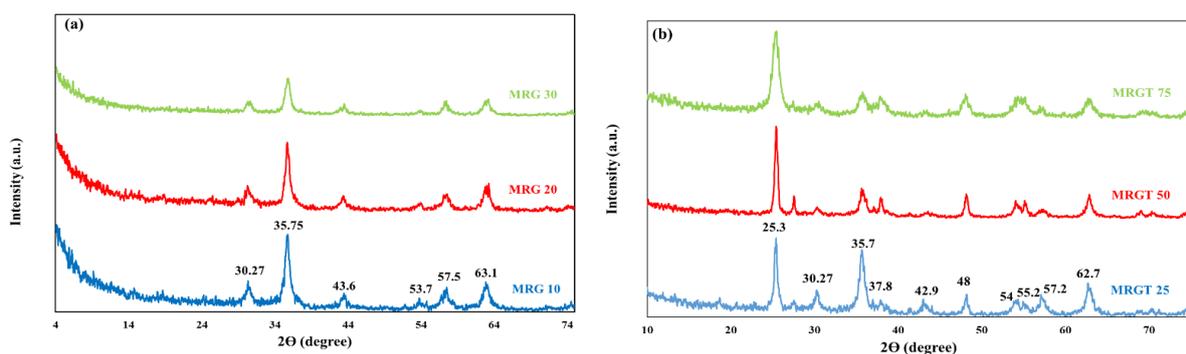


Fig. 1. X-ray diffraction patterns of (a) MRGx and (b) MRGTx samples

In fig (2.a) Raman spectroscopy is a technique that is broadly used to acquire basic data about carbon-based materials. The primary components in the Raman spectra of graphitic carbon-based materials are the G and D peaks. The normal features in Raman spectra were the D band situated at  $1345\text{cm}^{-1}$  and the G band at  $1592\text{cm}^{-1}$ . The G peak match to the optical  $E_{2g}$  phonons at the Brillouin zone center which outcome from the bond stretching of  $sp^2$  carbon

pairs in both, rings and chains. The D peak demonstrates breathing mode of aromatic rings which grow due to the defect in the sample. The D peak intensity is usually utilized as a measure of the level of disorder. The shift and shape of the overtone of the D peak, called as 2D peak around  $2680\text{ cm}^{-1}$ , can be related to the quantity of graphene layers (N) [31].

In Raman spectrum of MRGx (where x is 10, 20 and 30) there are three peaks situated at 217, 272,  $367\text{ cm}^{-1}$ , which are related to  $\text{Fe}_3\text{O}_4$  nanoparticles as show in Fig (2.b). The wide peaks showed at  $1357\text{ cm}^{-1}$  and  $1592\text{ cm}^{-1}$  are related to the D and G groups of reduced graphene oxide, respectively [28]. In MRG10, MRG 20 and MRG 30 the two distinctive peaks of G and D bands still coexist. This means that despite of reduction of GO's band intensity through loading of  $\text{Fe}_3\text{O}_4$ , there is no exfoliation between the layers of graphene happened. This affirming that it was a chemical bonding between  $\text{Fe}_3\text{O}_4$  and GO which coming from replace oxygen in GO by  $\text{Fe}_3\text{O}_4$  to save exfoliation between the Layers of graphene [32].

Raman spectrum of MRGTx samples presented in Fig (2.c), It display two strong peaks, meant as the disorder peak (D, situated at  $1357\text{ cm}^{-1}$ ) and the graphitic peak (G, at  $\sim 1592\text{ cm}^{-1}$ ) which are distinctive to reduced graphene oxide fingerprint.

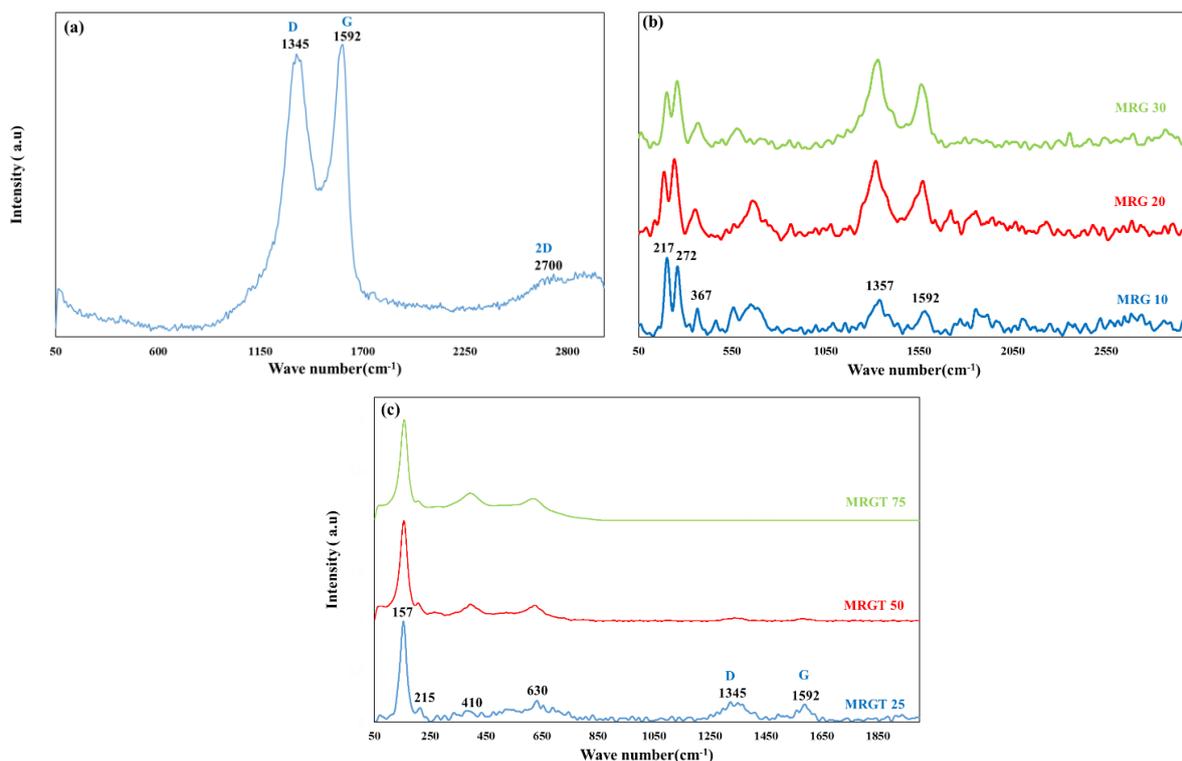


Fig. 2. Raman spectra of all prepared samples

For  $\text{TiO}_2$ , because of *E<sub>g</sub>* vibration modes of anatase, a sharp Raman scattering peak for free status anatase nanoparticles is usually seen at  $143\text{ cm}^{-1}$ . Conversely, the peak was seen at  $157\text{ cm}^{-1}$ . The blue shift from  $143\text{ cm}^{-1}$  to  $157\text{ cm}^{-1}$  might be ascribed to the interaction of RG and  $\text{TiO}_2$ . The peaks at  $\sim 215, 410, 520$  and  $630\text{ cm}^{-1}$  are related to the different vibration modes of anatase [33].

The morphologies of the magnetite with various amounts of reduced graphene oxide (10%, 20% and 30%) are seen by HRTEM in Fig (3- a, b and c respectively). As indicated the gray part demonstrate the RG nanoparticles, while the dark part demonstrates the aggregates of

$\text{Fe}_3\text{O}_4$  nanoparticles[34]. The large amount of  $\text{Fe}_3\text{O}_4$  is kept onto RG substrates, which shows that some  $\text{Fe}^{3+}$  ions could be reduced to  $\text{Fe}_3\text{O}_4$  by in-situ strategy [35]. The magnetite can be distinguished clearly in MRG (10 % and 20% respectively). Conversely at MRGO 30% , Graphene tend to cover whole the nanoparticles and hinder activity of these particles[36].

The HRTEM of the MRGT 25, 50 and 75 composites was observed in Fig (3- g, h and i respectively). The RG sheets were covered with a  $\text{TiO}_2$  layer, demonstrating loading amount of  $\text{TiO}_2$ , which is combatable with the EDX results. The edge of graphene can be clearly recognized, persuading us the existence of MRG layer under the  $\text{TiO}_2$  [37].

The EDX spectrum shows Fe, O and C peaks which are related to MRG composite, (the O peak coming form  $\text{Fe}_3\text{O}_4$ ). Thus in general there is a good contact between  $\text{Fe}_3\text{O}_4$  nanoparticles and RG [38]. Moreover that gradation of amount Fe refers to well synthesized amount between magnetite and graphene this appear in Fig (3- d, e and f). Also The EDX results for MRGTx composites refer to the existence of Ti, Fe, O and C elements with prepared percent show in Fig (3- j, k and l), which confirm the existence of  $\text{TiO}_2$  on the surface of MRGO [39].

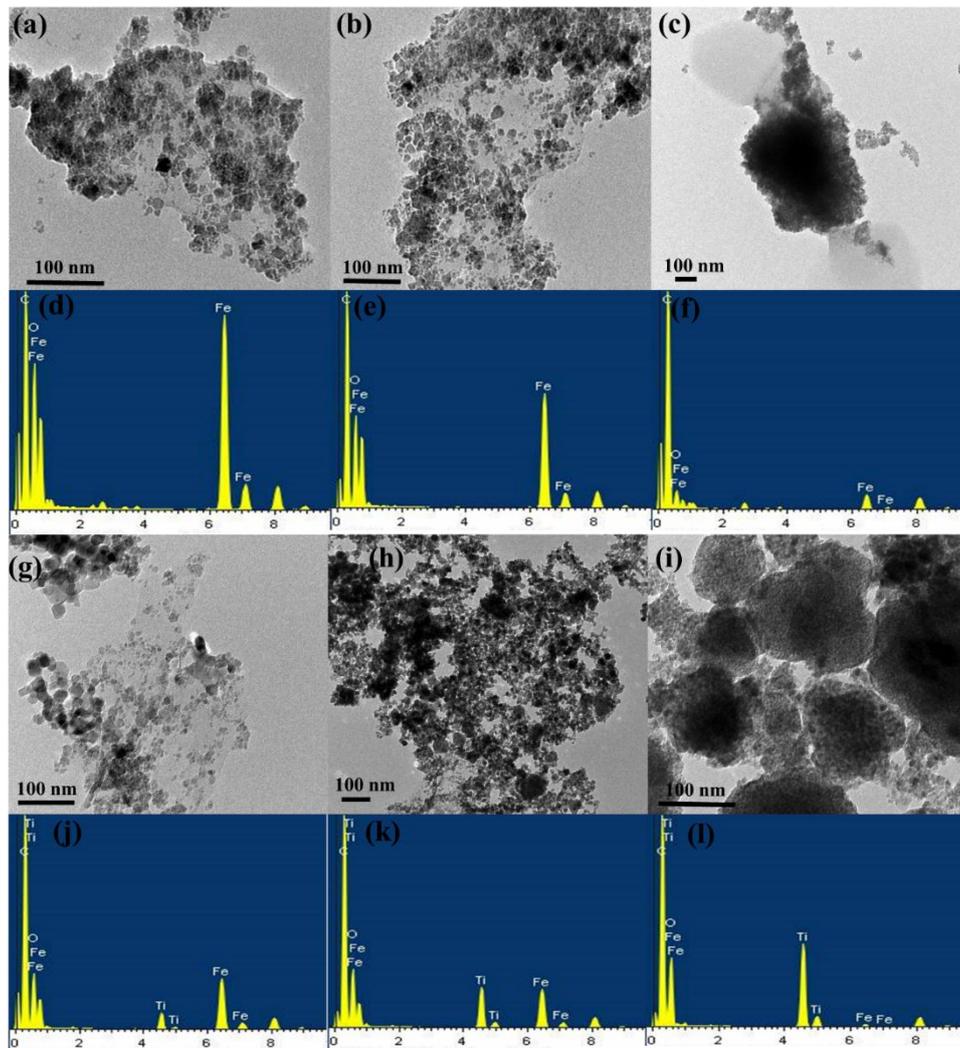


Fig. 3. HRTEM and EDX of (a) and (d) for MRG 10, (b) and (e) for MRG 20, (c) and (f) for MRG 30, (g) and (j) for MRGT 25, (h) and (k) for MRGT 50, (i) and (l) for MRGT 75

In FT-IR spectra of MRG and MRGT show in Fig (4 a and b), the broad peak located at about  $3400\text{ cm}^{-1}$  which was assigned to the stretching vibration of O-H. The absorption Peak at  $600\text{ cm}^{-1}$  can be considered as a combination of Fe–O vibrations, demonstrating the effective blending of  $\text{Fe}_3\text{O}_4$  nanoparticles and RGO sheets[40]. The absorption band showing up at  $1570\text{ cm}^{-1}$  is related to the skeletal vibration of RGO sheets, which give data about the formation of RGO [41].

The peaks at  $1400\text{ cm}^{-1}$  appear due to the Ti–O–C vibration, which confirm the successful interaction between Ti and C and, therefore the impregnation of  $\text{TiO}_2$  on reduced graphene oxide (RG) while the absorption band at  $584\text{ cm}^{-1}$  give information about the presence of Ti–O–Ti bond on the surface of MRGT nanocomposite [42].

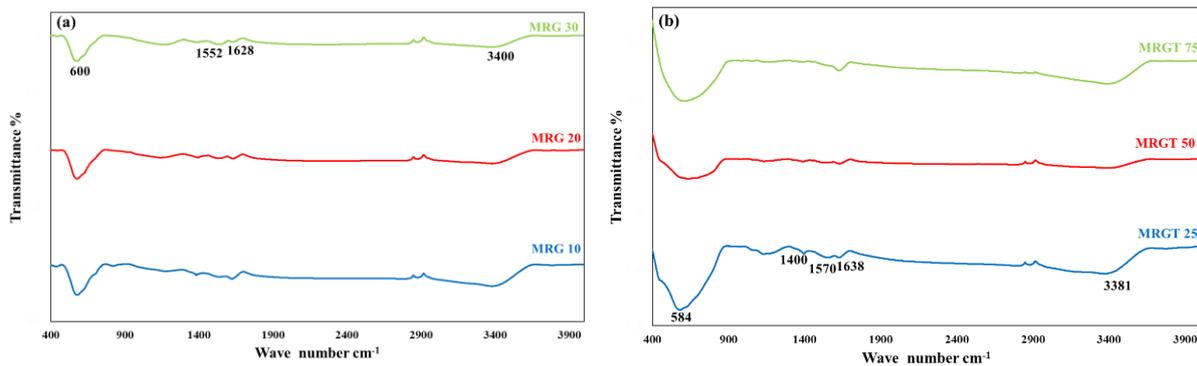


Fig. 4. FT-IR spectra of MRG and MRGT samples

VSM technique was done to know the magnetic behavior of both  $\text{Fe}_3\text{O}_4$  nanopowder and MRG nanocomposite. VSM find the magnetic momentum by vibrating the sample in perpendicular to the regular magnetic field, which is extremely normal and suitable technique for find the magnetic behavior .VSM was done at room temperature and the magnetic behavior results of MRG 10, 20 and 30 are shown in the Fig (5-a). The  $\text{Fe}_3\text{O}_4$  nanoparticles and MRG nanocomposites demonstrate magnetization hysteresis loops and these are S-like curves. All samples of MRG 10, 20 and 30 demonstrated that they have a super paramagnetic behavior with a saturation magnetization of 50, 47.3 and 40.6 emu/g respectively, which is great for separating nanocomposites from the solution mixture directly and use it again for further uses.  $\text{Fe}_3\text{O}_4$  nanopowder displays ferromagnetic behavior with a saturation magnetization of 51 emu/g, which was observed to be higher than that of MRG nanocomposites. The less value of saturation magnetization in MRG nanocomposites was because of the existence of RG nanosheet in the nanocomposites. [43], [44].

According to the magnetization curves in Fig (5-b), the saturation magnetization of the MRGT (25,75and 75% proportion) decrease to 34.7, 22.4 and 9.4 emu/g respectively from 51 emu/g in the primary  $\text{Fe}_3\text{O}_4$  sample because of the presence of nonmagnetic materials in the surface of MRG nanocomposite, which is still adequate for rapidly magnetic separation with a suitable magnet field. However , the reversibility in hysteresis loop affirms that no that no aggregation occurs to the nanoparticles in the magnetic fields .[45]

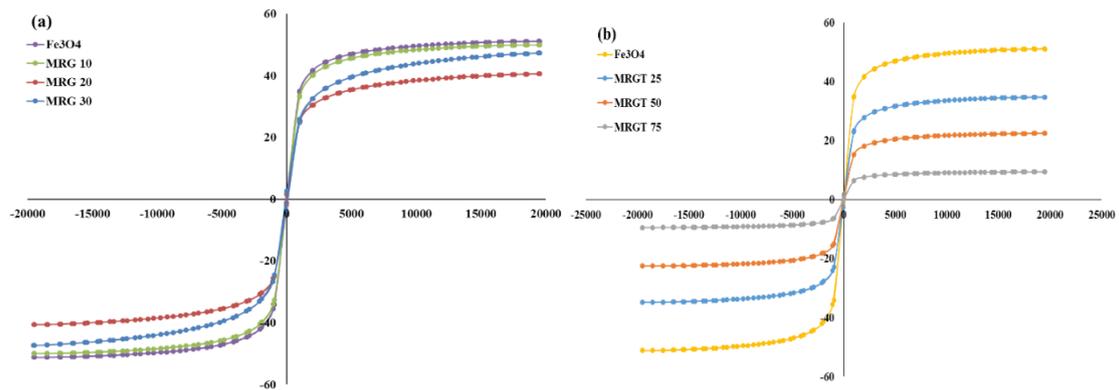


Fig. 5. VSM measurements for magnetic nanoparticles, (a) for MRGx samples and (b) for MRGTx samples.

Fig (6-a) Presents UV-vis diffuse reflectance spectra of MRG 10, 20 and 30 samples. The spectra of sample MRG 20 have a red shift for about 25 nm relative to MRG 10. This indicates the decrease of the band gap energy. The  $E_g$  values were calculated from the  $(F(R)h\nu)^{1/2}$  versus  $h\nu$  plots, where  $F(R)=(1-R)/2R$  (Todorova et. al. 2008). The absorption edges were found to be 598 nm, this value corresponds to the band gap energy of  $E_g=1.8\text{eV}$  for MRG 20. Then little blue shift for MRG 30, to give the  $E_g$  value 1.6 eV as show in Fig (6-b) this refers to blocking of magnetite by enriched of reduced graphene layers. So the best ratio of RGO on  $\text{Fe}_3\text{O}_4$  is 20% is given smallest  $E_g$  in range semiconductor.

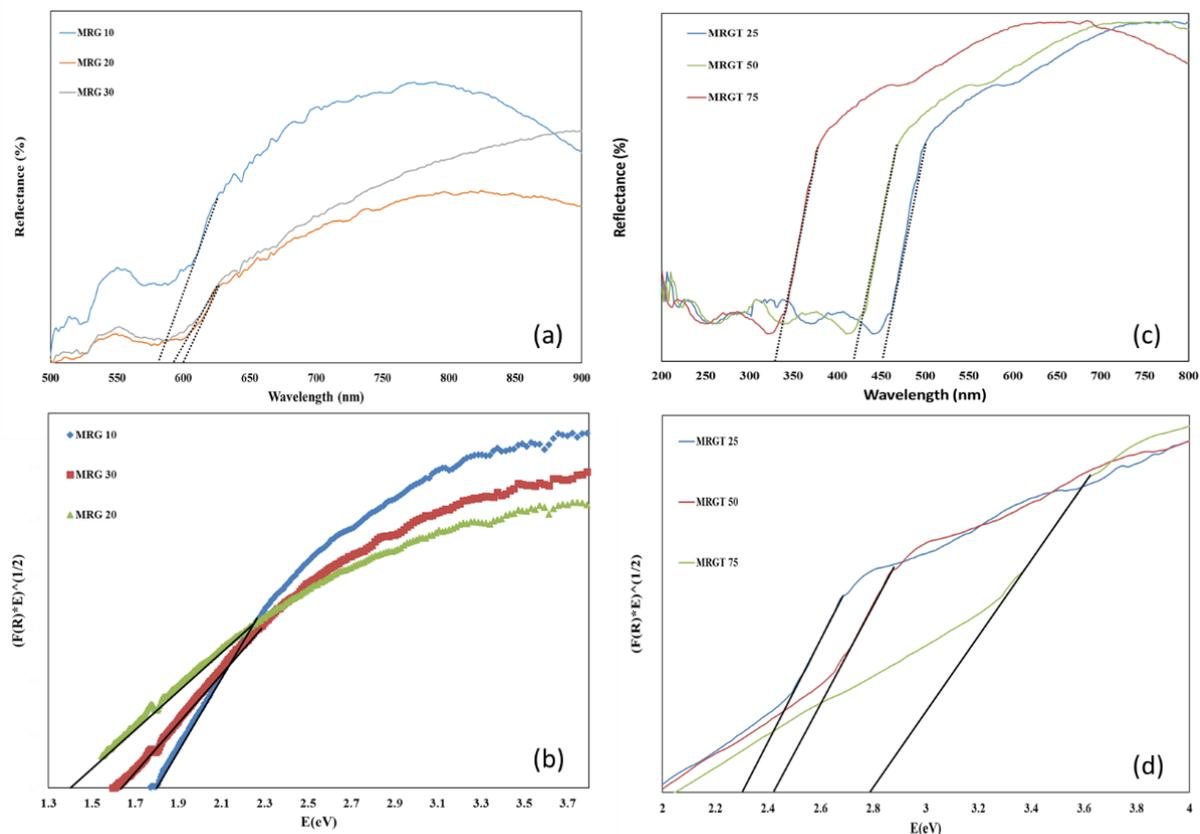


Fig. 6. Diffuse reflectance and  $E_g$  band gap for all prepared samples

### 3.2 Catalytic activity of MRGx and MRGTx samples

After added TiO<sub>2</sub> on MRG20 by different ratio presented UV–vis diffuse reflectance spectra in Fig (6-c). From MRGT 25, 50 and 75 given blue shift where at MRGT75 has band gap at 2.8 eV as shown in Fig (6-d) and this is closely to an ultra violet region. But MRGT 50 has E<sub>g</sub> 2.4 eV in the visible region, hence we obtained all benefits of high efficiency from TiO<sub>2</sub> and enhanced in visible region by an optimum ratio of magnetite reduced graphene.

In order to determine the optimal photocatalyst of MRG (either MRG10, 20 or 30) to use it as a substrate to loading TiO<sub>2</sub> nanoparticle on its surface, an experiment was performed at room temperature by varying the time from 0 to 210 min. The effect of time on the rate of degradation indicates that optimum catalyst loading was MRG 20 as shown in Fig (7). The rate of degradation increases with increase the amount of loading graphene on the surface of magnetite from MRG10 to MRG 20. In MRG 30 we observed that the rate of degradation is decreased, This is probably due to increasing in the amount of loading graphene (RG) on magnetite surface is lead to the complete cover of nanoparticles and hinder the photocatalytic activity of this particles this matched with data from band gap and HR-TEM.

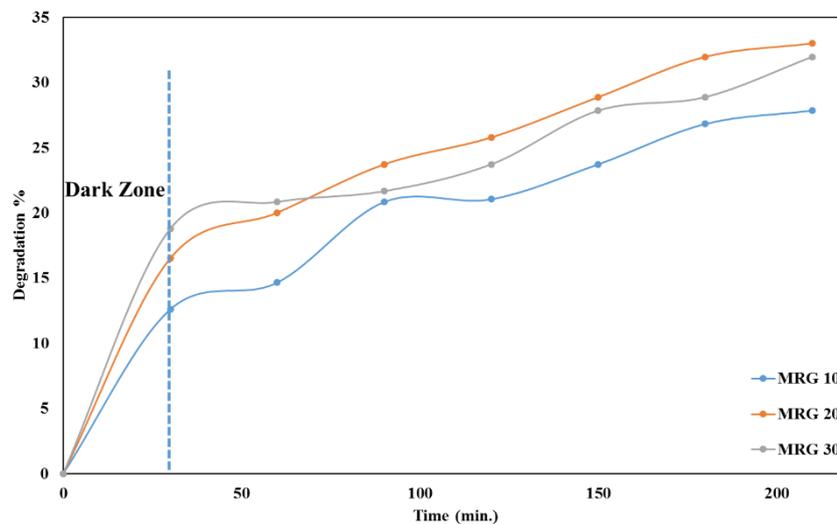


Fig. 7. Photocatalytic degradation of tartrazine by MRGx samples

Study of TZ degradation by MRGT 25, 50 and 75, an experiment was performed at room temperature by varying the time from 0 to 210 min. The effect of time on the rate of degradation indicates that the optimum catalyst loading on MRG 20 is MRGT 50 as shown in Fig (8). The rate of degradation increases with increase the amount of loading TiO<sub>2</sub> on the surface of MRG 20 from MRGT 25 to MRGT 50. In MRGT 75 we observed that the rate of degradation is decreased, This is probably due to increasing in the amount of loading TiO<sub>2</sub> (T) on MRG 20 surface where the excess of titanium dioxide have shifted the composite to work in UV region not in the visible region, so give less activity. And by comparing with TiO<sub>2</sub> which give activity less than 10% of degradation this refer to not working of TiO<sub>2</sub> in visible light and also strong TZ day to degradation by it only.

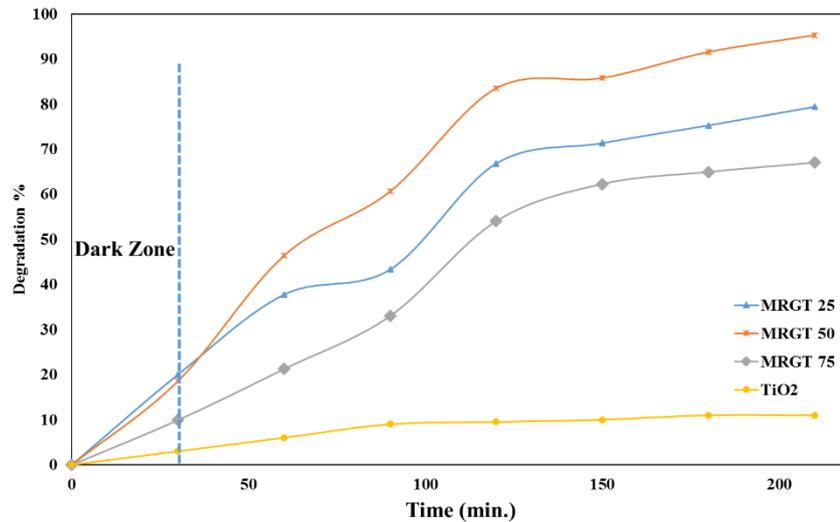


Fig. 8. Photocatalytic degradation of TZ by MRGT25, 50, 75 and TiO<sub>2</sub> samples

### 3.3 MRGT50 effective conc. and its endurance against high concentrations

To determine the optimal amount of the MRGT50 sample. The experiments are performed at room temperature by varying the amount of catalyst from 0.15 to 0.3 g/l. The effect of MRGT 50 loadings on the rate of degradation indicates that optimum catalyst concentration is 0.2g/l show in Fig (9-a). The rate increases with increase in catalyst concentration from 0.15 to 0.2g/l. This is probably due to increasing in the amount of photocatalyst, which increases the number of photons and dye molecules absorbed. The increase of the catalyst concentration to more than 0.2g/l results in the decrease of degradation rate. This phenomenon may be explained by aggregation of MRGT 50 at high concentrations. It causes a decrease in the number of surface active sites, which increases the opacity and lights scattering of MRGT 50, leading to a decrease in the passage of irradiation through the sample.

Experiments are performed at a different concentration of TZ 50, 100 and 200 ppm, against MRGT 50 [0.2g/l], at room temperature. Degradation rate appears to decrease with increasing TZ concentration in Fig (9-b). This is because active sites of the photocatalyst are occupied by TZ molecules and fewer photons can reach the surface of the catalyst. Therefore, the possibility of photocatalyst excitation is diminished. At higher concentration levels, the screening effect also dominates and hence, degradation efficiency decreases.

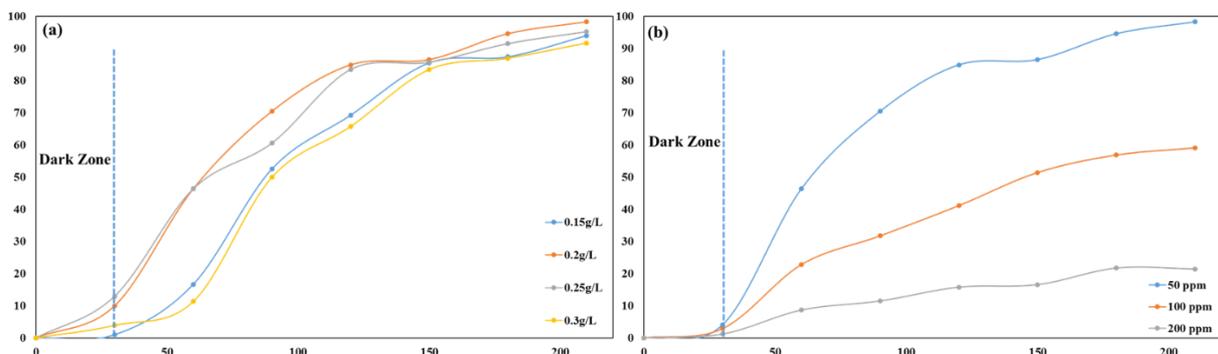


Fig. 9. (a) Effect of weight catalyst MRGT50 (b) Effect of TZ concentration

## 4. Conclusion

A magnetically separable MRG and MRGT photocatalyst has been successfully prepared. TEM observations indicate that graphene sheets are fully exfoliated and decorated with Fe<sub>3</sub>O<sub>4</sub> and TiO<sub>2</sub> nanoparticles. The photocatalytic activity measurements show coupling Fe<sub>3</sub>O<sub>4</sub> and TiO<sub>2</sub> nanoparticles with graphene sheets which lead to high photocatalytic activity in degradation of TZ under visible light irradiation reach to more than 95% degradation of pollutant. In contrast, TiO<sub>2</sub> alone is almost inactive in our experimental condition. Easy separation of photocatalyst by a magnet to prevent secondary pollution.

## 5. References

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